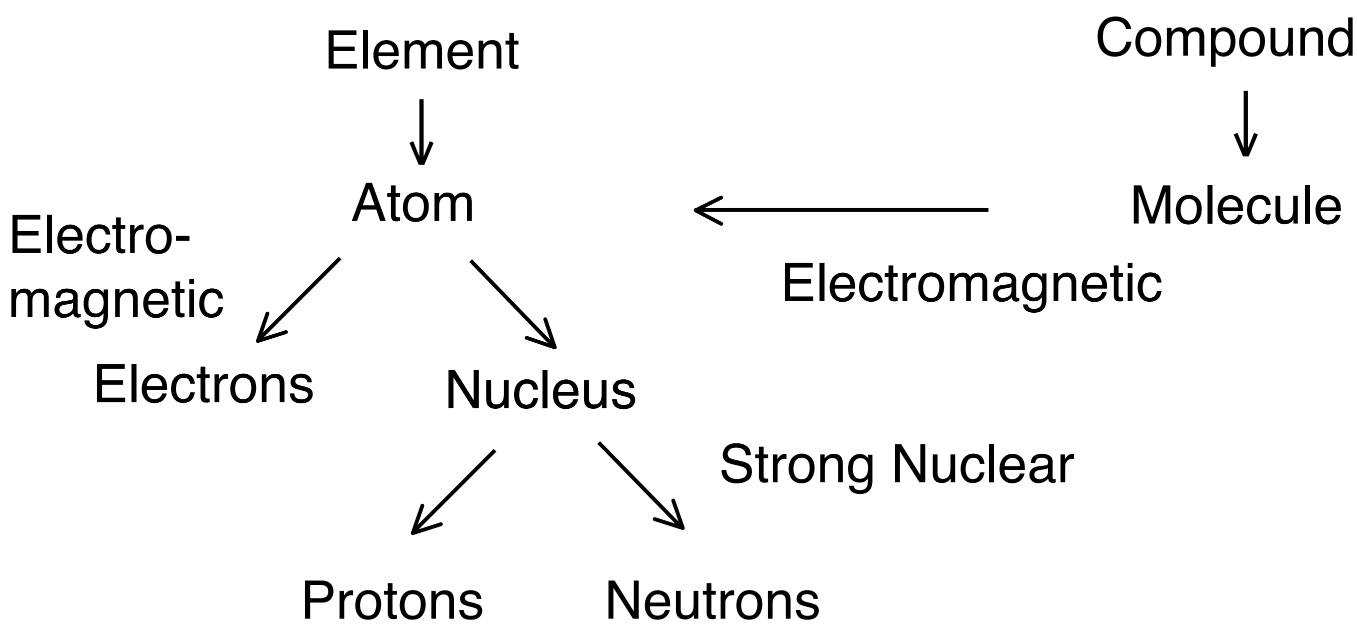
Cosmic Evolution, Part II

Heavy Elements to Molecules

Heavy elements → molecules

First a review of terminology:



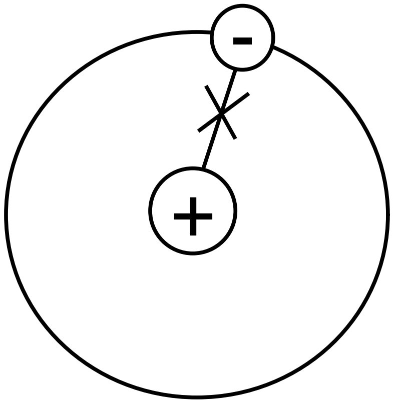
Neutral atom: # Electrons = # protons

ion: ≠

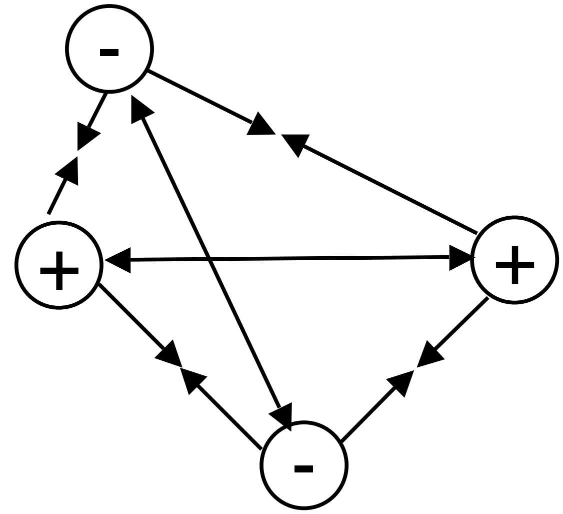
e.g. C^{+2} Carbon nucleus + 4 (6-2) electrons

Forces

H atom



H_2
Molecule



Attractive   Repulsive

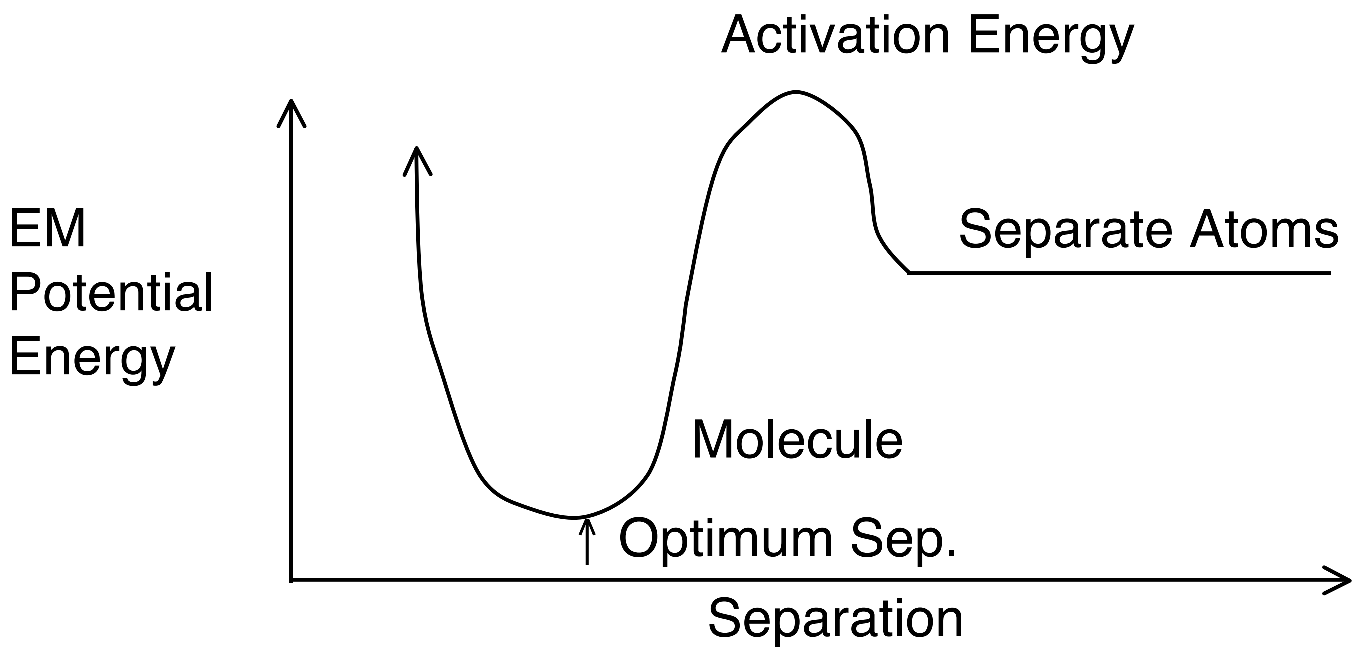
Molecule: Repulsive ~ Attractive

More delicate than atoms,
can be <u>much</u> more complex

“Bond” is sharing of electrons

Is molecule stable?

Yes, if EM potential energy less than separate atoms



Activation energy lower → T ~ 100 - 1000 K

(Room Temperature)

Questions

- Why is room temperature around 300 K?
- How commonly is this temperature found in the Universe?

Conventions:

H_2 $\quad$ H – H (Bond) $\quad$ CO_2 $\quad$ O = C = O (Double Bonds)

Maximum # of Bonds:	H	1
	O	2
	N	3
	C	4

Carbon very versatile
→ Complex chemistry

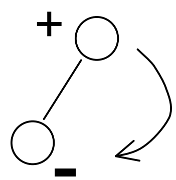
Interstellar Molecules

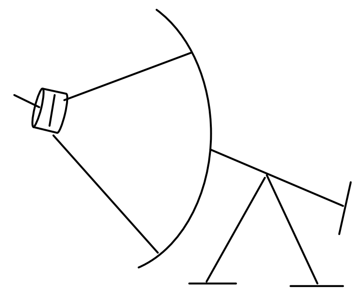
Exist as gas (individual molecules)

A few known in 1930's

Many more since 1968 - Radio astronomy

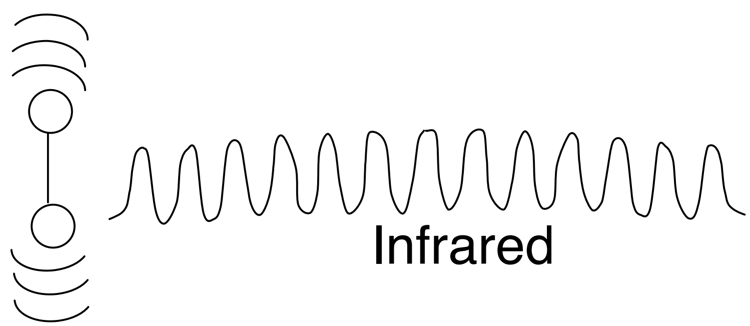
Rotation

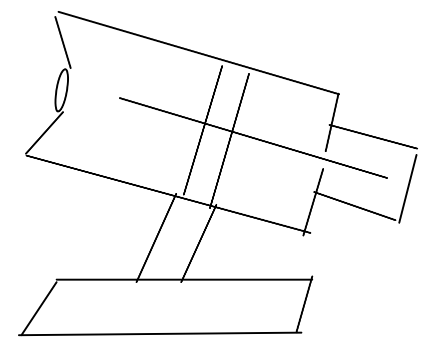




Radio Telescope

Vibration



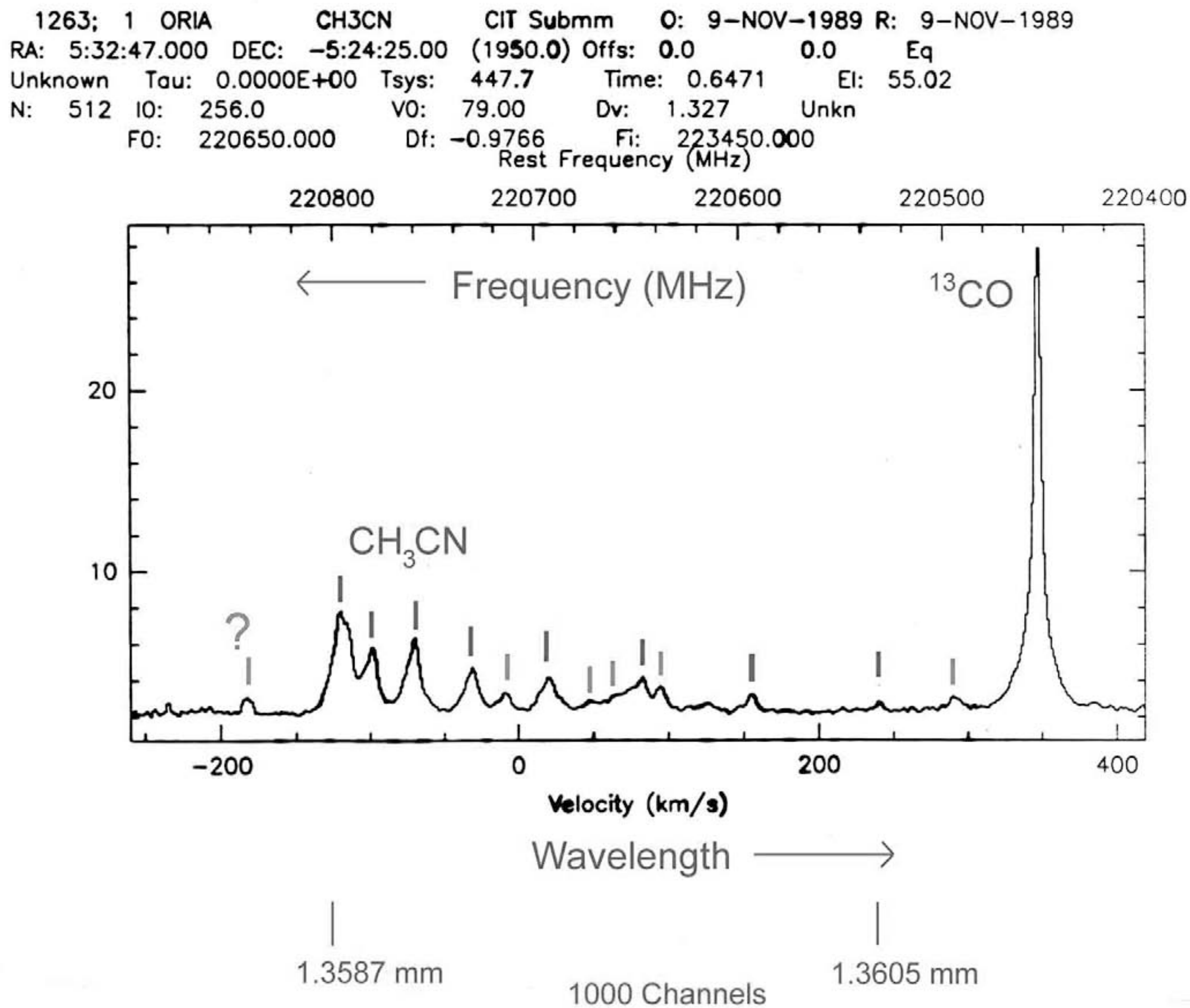


Optical Telescope

How we detect Interstellar Molecules

Radio Spectroscopy (Mostly $\lambda \sim$ 1- 3 mm)

+ Precise knowledge of wavelengths for different molecules



Appendix 2

Interstellar Molecules

Molecular Ions: HCO^+, HCS^+, HOC^+, N_2H^+

Species	Name
H_2	molecular hydrogen
C_2	diatomic carbon
CH	methylidyne
CH^+	methylidyne ion
CN	cyanogen
CO	carbon monoxide
CO^+	carbon monoxide ion
CS	carbon monosulfide
OH	hydroxyl
HCl	hydrogen chloride
NH	
NO	nitric oxide
NS	nitrogen sulfide
SiC	silicon carbide*
SiO	silicon monoxide
SiS	silicon sulfide
SiN	silicon nitride
SO	sulfur monoxide
PN	
CP	*
SO^+	sulfoxide ion
NaCl	sodium chloride*
AlCl	aluminum chloride*
KCl	potassium chloride*
AlF	aluminum fluoride*†
FeO	iron monoxide
HF	
SH	
H_3^+	protonated hydrogen
C_2H	ethynyl
CH_2	methylene †
HCN	hydrogen cyanide
HNC	hydrogen isocyanide
HCO	formyl
HCO^+	formyl ion
HCS^+	thioformyl ion
HOC^+	isoformyl ion †
N_2H^+	protonated nitrogen
HNO	nitroxyl
H_2O	water
H_2S	hydrogen sulfide
H_2N	hydrogen nitride
N_2O	nitrous oxide

Species	Name
CO_2	carbon dioxide
OCS	carbonyl sulfide
SO_2	sulfur dioxide
SiC_2	silicon dicarbide*
SiCN	
AlCN	
C_2S	
C_2O	dicarbon monoxide †
C_3	triatomic carbon*
MgCN	magnesium cyanide*
MgNC	magnesium isocyanide*
NaCN	sodium cyanide*
C_2H_2	acetylene
C_3H	propynylidyne (l and c)
H_2CO	formaldehyde
H_2CN	
HC_2N	
NH_3	ammonia
HNCO	isocyanic acid
$HOCO^+$	
$HCNH^+$	
HNCS	isothiocyanic acid
C_3N	cyanoethynyl
C_3O	tricarbon monoxide
C_3S	
H_2CS	thioformaldehyde
H_3O^+	hydronium ion
SiC_3	
C_4H	butadiynyl
C_3H_2	cyclopropenylidene
H_2CCC	propadienylidene
HCOOH	formic acid
CH_2CO	ketene
HC_3N	cyanoacetylene
HNC_3	
CH_2CN	cyanomethyl
NH_2CN	cyanamide
CH_2NH	methanimine
HC_2NC	
CH_4	methane

Species	Name
H_2COH^+	protonated formaldehyde
SiH_4	silane*
C_4Si	*
C_5	pentatomic carbon*
C_5H	pentynylidyne
C_5N	
C_2H_4	ethylene*
H_2CCCC	butatrienylidene
CH_3OH	methanol
CH_3CN	methyl cyanide
CH_3NC	methyl isocyanide
CH_3SH	methyl mercaptan
NH_2CHO	formamide
HC_3HO	propynal
HC_3NH^+	
C_6H	
CH_2CHCN	vinyl cyanide
CH_3C_2H	methylacetylene
CH_3CHO	acetaldehyde
CH_3NH_2	methylamine
C_2H_4O	ethylene oxide
CH_2CHOH	vinyl alcohol

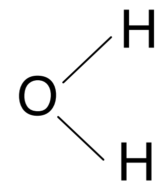
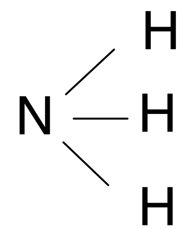
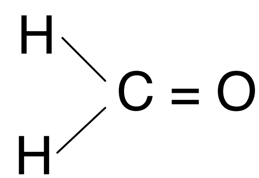
Species	Name
HC_5N	cyanodiacetylene
C_7H	
$HCOOCH_3$	methyl formate
CH_3C_3N	methylcyanoacetylene
CH_3COOH	acetic acid
H_2C_6	
CH_2OHCHO	glycolaldehyde
CH_3C_4H	methyldiacetylene
CH_3CH_3O	dimethyl ether
CH_3CH_2CN	ethyl cyanide
CH_3CH_2OH	ethanol
HC_7N	cyanohexatriyne
C_8H	
CH_3C_4CN	†
CH_3CH_3CO	acetone
NH_2CH_2COOH	glycine†
CH_2OHCH_2OH	ethylene glycol
HC_9N	cyano-octa-tetra-yne
$HC_{11}N$	cyano-deca-penta-yne

* Detected in circumstellar envelopes only
† tentative

— Important Probe of conditions
— Discovered in Infrared
— Discovered in UV
— Relevant to the Origin of Life

Look at Appendix 2

Important Examples:

Water	H_2O	
Ammonia	NH_3	
Formaldehyde	H_2CO	

Others of Note: CO Most common after H_2

HCN, HC_3N, … $HC_{11}N$ → Carbon chains

CH_4 (Methane)

PAHs (Polycyclic aromatic hydrocarbons)

<u>3 Lessons</u>

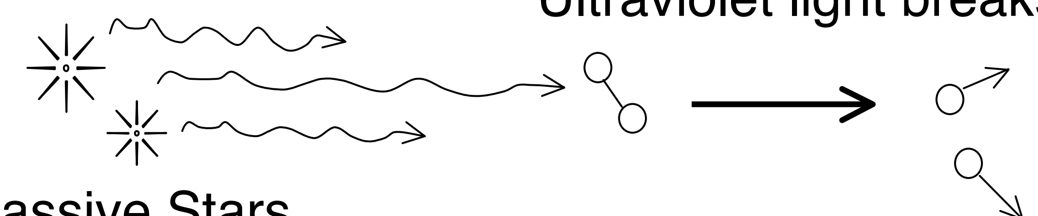
1. Complexity (Up to 13 - atoms) is extraterrestrial
   May be more complex (Hard to detect)
   Glycine ? 1994
   Polycyclic Aromatic Hydrocarbons (PAHs)
   (Infrared evidence)

2. Dominance of Carbon
   Carbon Chemistry not peculiar to Earth

3. Formation & Destruction   <u>Analogous</u> to early Earth

Destruction:

Ultraviolet light breaks bonds



Massive Stars

Protection by dust grains: scatter and absorb ultraviolet

Dust

Studies of how they scatter and absorb light
(Ultraviolet → Visible → Infrared)

⇒ Two types, range of sizes up to 10^{-6} m

Carbon
PAHs → Graphite
~ Soot

Silicates
Si + O + Mg, Fe, …

Both Produced by old stars

Formation of Interstellar Molecules

1. H_2

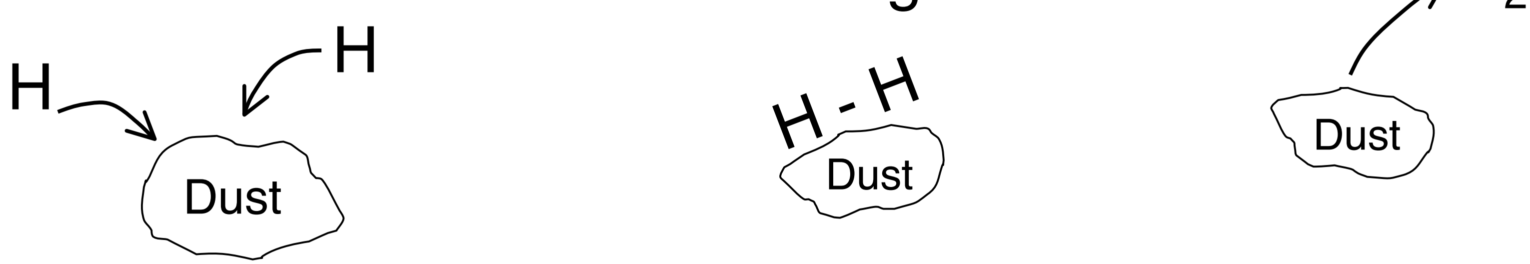
   Must lose the potential energy difference before it falls apart (~ 10^{-14} s)

   Collisions: OK in lab, too slow in space

   Emit photon: <u>very</u> slow for H_2 (10^7 s)

   $H + H + \text{catalyst} = H_2 + \text{catalyst}$

   ↑ surface of dust grain



Formation of Interstellar Molecules

2. More complex molecules

Problem is activation energy barrier

$T \sim 10$ K << Barrier

Use reactions **without** activation energies

e.g. Molecular ions, like HCO^+

Cosmic Ray

$$\xrightarrow{\hspace{4em}} H_2 \rightarrow H_2^+$$

$$H_2^+ + H_2 \rightarrow H_3^+ + H$$

$$H_3^+ + CO \rightarrow HCO^+ + H_2$$

$$XH^+ + e^- \rightarrow X + H$$

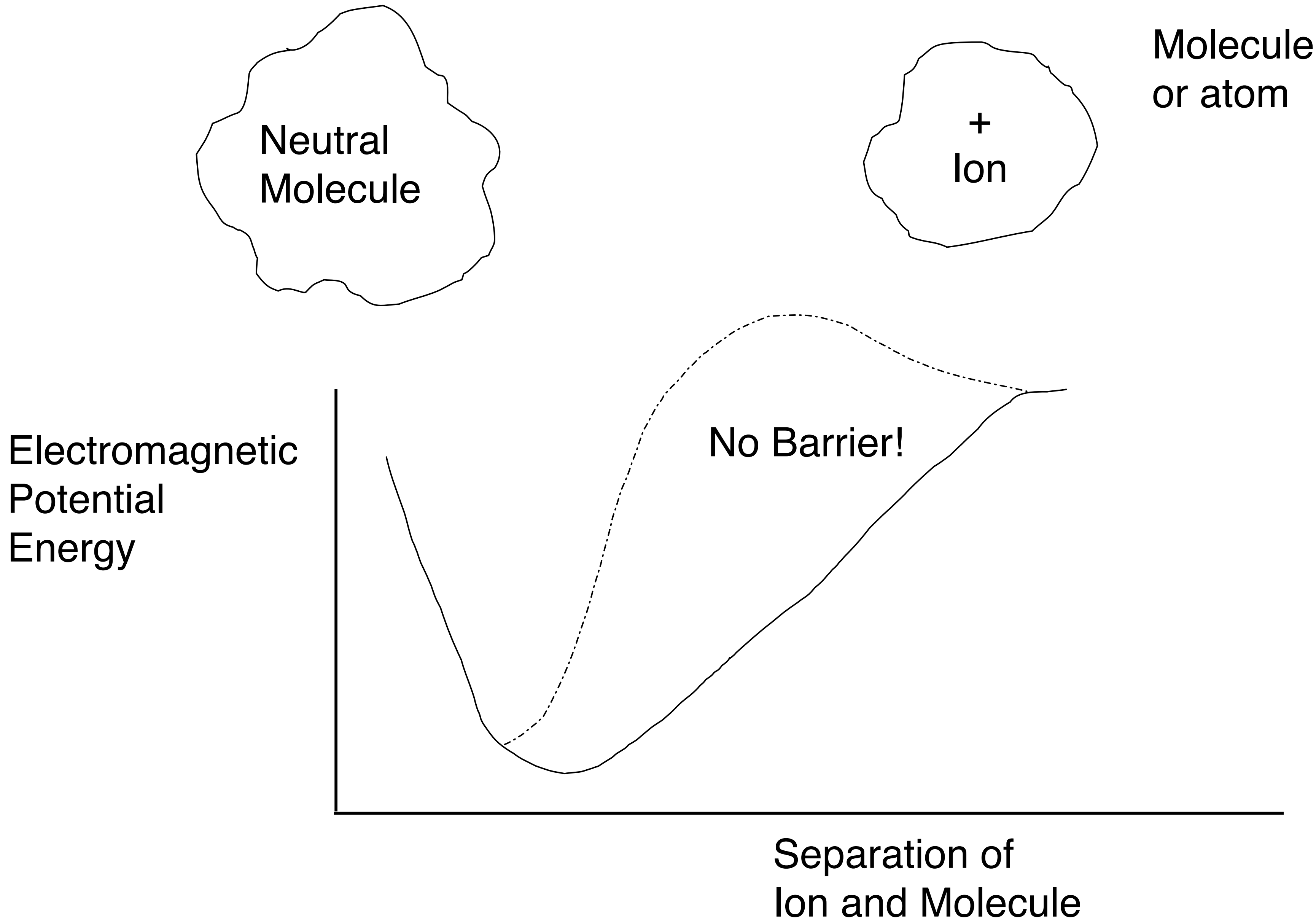
Energy + simple mol.

$\rightarrow$ Reactive mol.

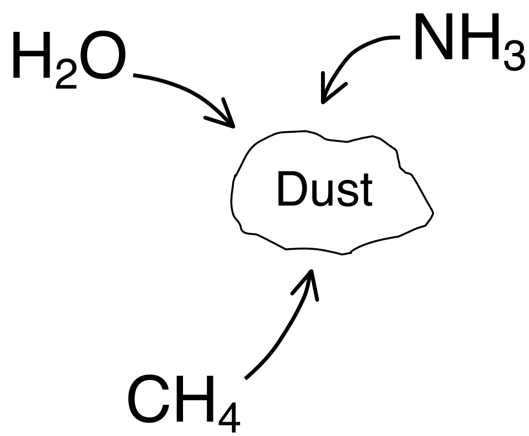
$\downarrow$

More complex

Ion - Molecule Reactions



Molecules on Dust Grains



Stick on grains
“ice”

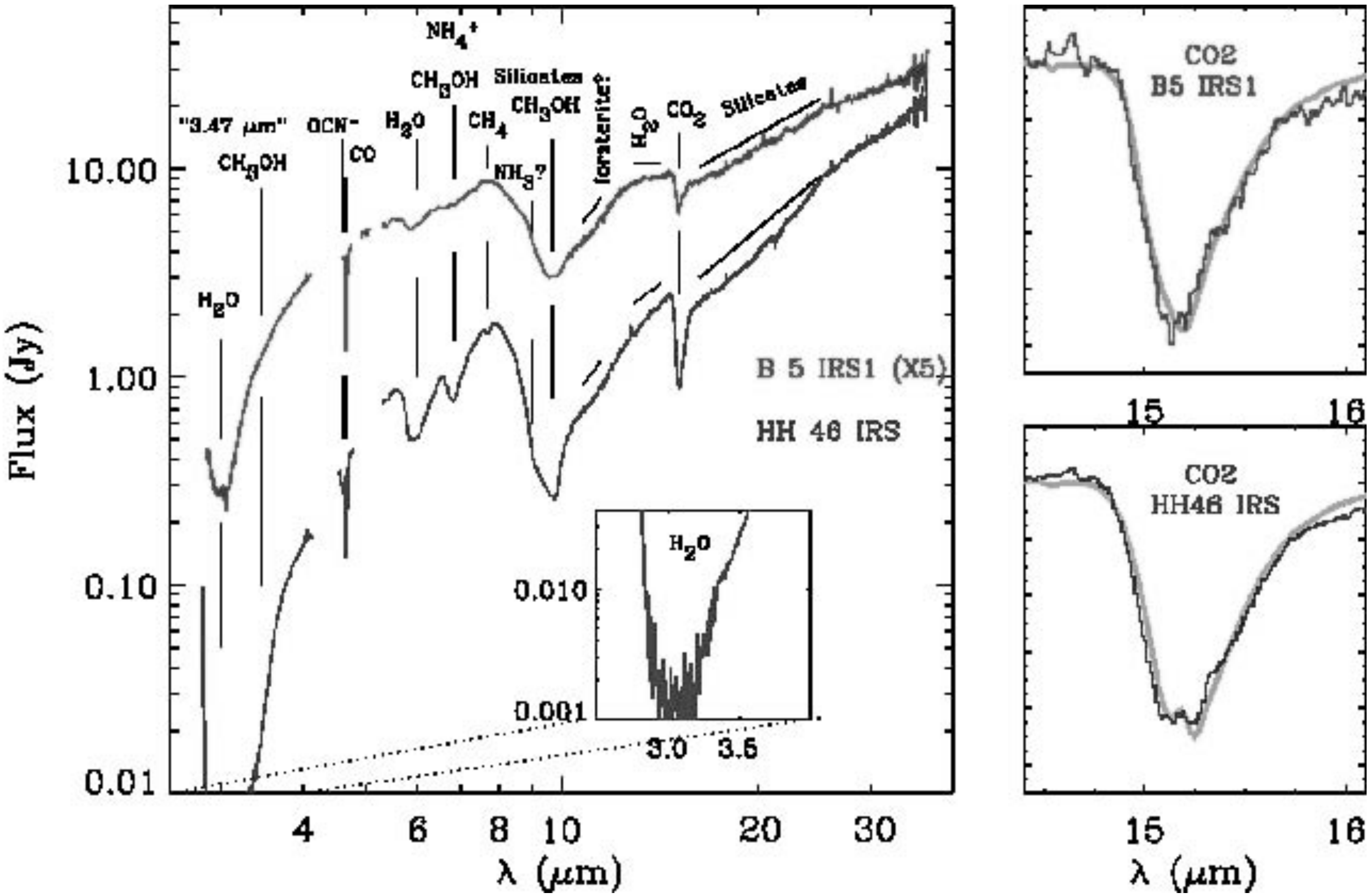
Infrared observations show this: as molecules

Vibrate, absorb infrared

e.g. H_2O absorbs at 3×10^{-6} m

CH_4 absorbs at 8×10^{-6} m

Ices on Dust Grains



Molecules on Dust Grains

Icy "mantles" contain H, O, C, N

Further reactions possible → more complex molecules (e.g. Ethanol)

→ Building blocks of life ?

→ Life ??? Hoyle and Wickramasinghe

New stars and planets form in same regions

Implications

1. Similar (Carbon-Dominated) Chemistry

2. Direct Role in Origin of Life?

3. Formation + Destruction
   analogous to Early Earth

Roles of Dust

1. Protection from UV
2. H_2 Formation
3. Depletion $\rightarrow$ Mantles of Ice
   H_2O, NH_3, CH_4, CO_2, HCOOH, …

   ↑ Methane

Star Formation

First factor in Drake Equation: The rate of star formation

Estimate of Average Star Formation Rate (R_*)

$$R_* = \frac{\text{\# of stars in galaxy}}{\text{lifetime of galaxy}} = \frac{N_*}{t_{gal}}$$

N_* : Count them? No

Use Gravity (Newton's Laws)

Sun orbiting center of galaxy at 250 km s^{-1} (155 miles per second)

Kinetic energy = $\frac{1}{2}$ gravitational potential energy

$$\frac{1}{2} M_\odot v^2 = \frac{1}{2} \frac{G M_g M_\odot}{R_g}$$

R_g ← Distance of Sun from center of galaxy

$$\frac{R_g v^2}{G} = M_g$$

Estimate of Average Star Formation Rate (R_*)

(R_g = 25,000 ly) $\rightarrow$ $M_g = 1.0 \times 10^{11}\, M_\odot$

Add stars outside Sun's orbit $\rightarrow$ $M_g \simeq 1.6 \times 10^{11}\, M_\odot$

$$N_* \simeq \frac{M_g}{\text{Avg. mass of star}} = \frac{1.6 \times 10^{11}}{0.4} = 4 \times 10^{11}$$

$t_{gal} \simeq 10^{10}$ yr (studies of old stars)

$$R_* \simeq \frac{4 \times 10^{11}}{10^{10}} \text{ stars} = 40 \text{ stars per year} \quad (5 - 50)$$

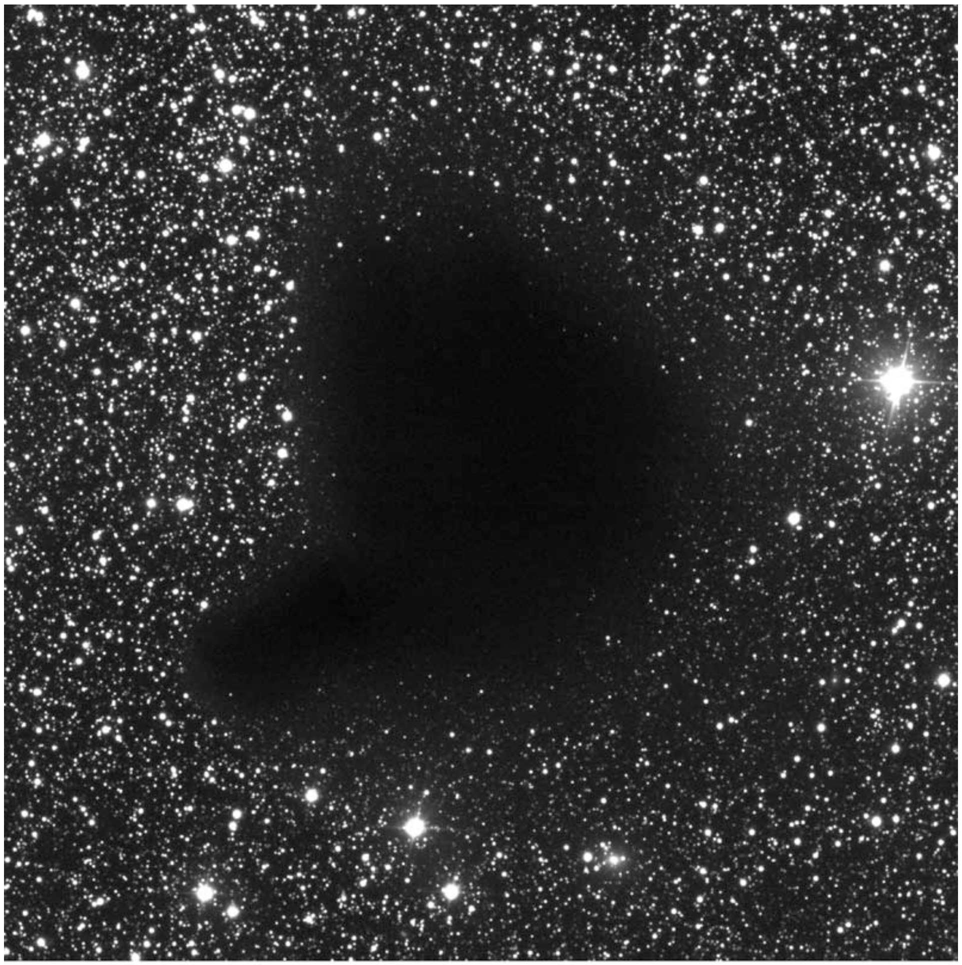
Star Formation

Current Star Formation

Molecular Clouds

- Composition
  - H_2 (93%), He (6%)
  - Dust and other molecules (~1% by mass)
    - CO next most common after H_2, He
- Temperature about 10 K
- Density (particles per cubic cm)
  - ~100 cm^{-3} to 10^6 cm^{-3}
  - Air has about 10^{19} cm^{-3}
  - Water about 3 x 10^{22} cm^{-3}
- Size 1-300 ly
- Mass 1 to 10^6 M_{sun}

A Small Molecular Cloud



The "Black Cloud" B68
(VLT ANTU + FORS1)

ESO PR Photo 20a/99 (30 April 1999)

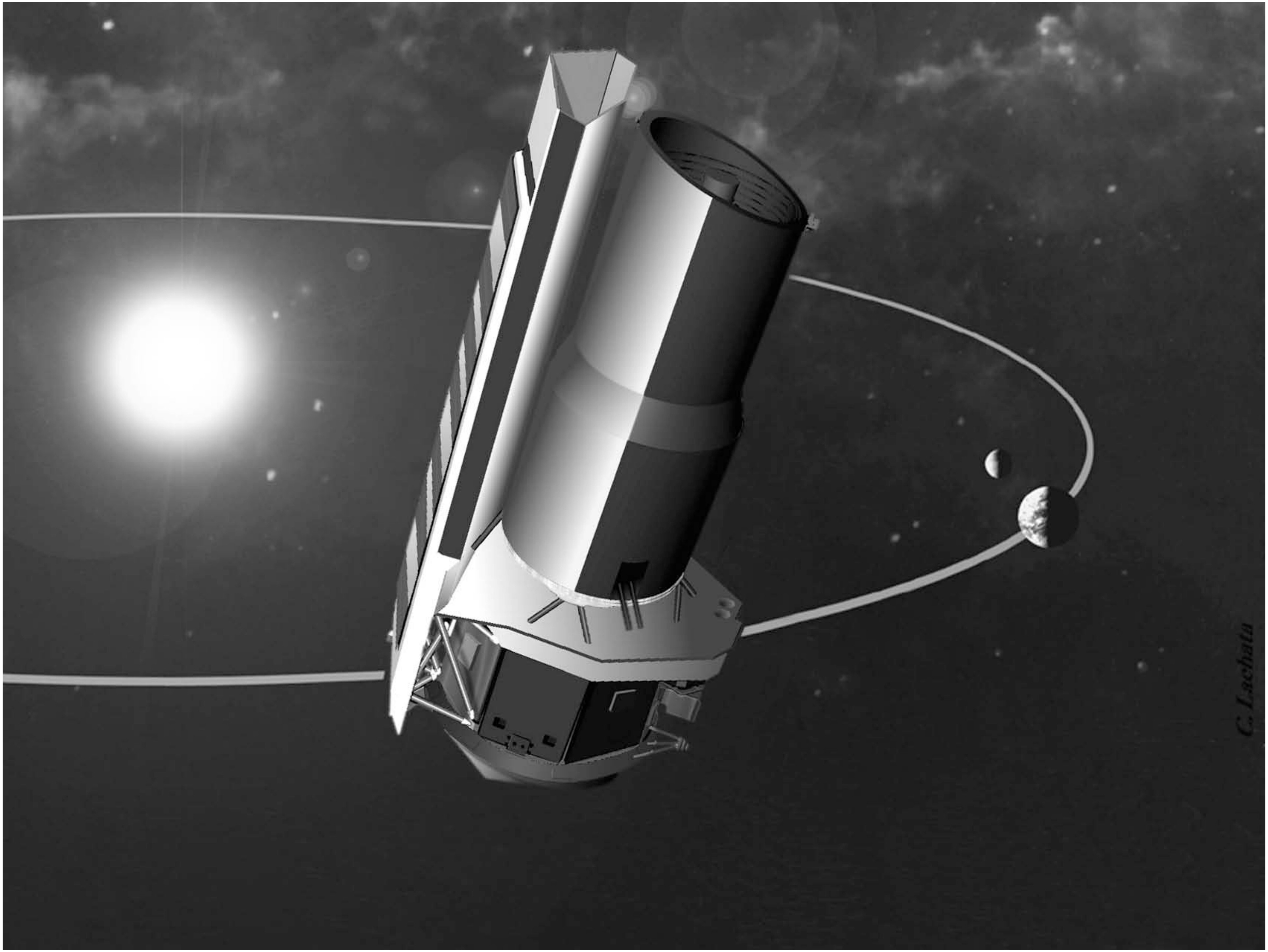
© European Southern Observatory

Current Star Formation

- Occurs in gas with heavy elements
  - Molecules and dust keep gas cool
  - Radiate energy released by collapse
  - Stars of lower mass can form
  - Mass needed for collapse increases with T
- Star formation is ongoing in our Galaxy
  - Massive stars are short-lived
  - Star formation observed in infrared

The Launch of The Spitzer Space Telescope

Spitzer Space Telescope Launched Aug. 2003, expect a 5 yr life.

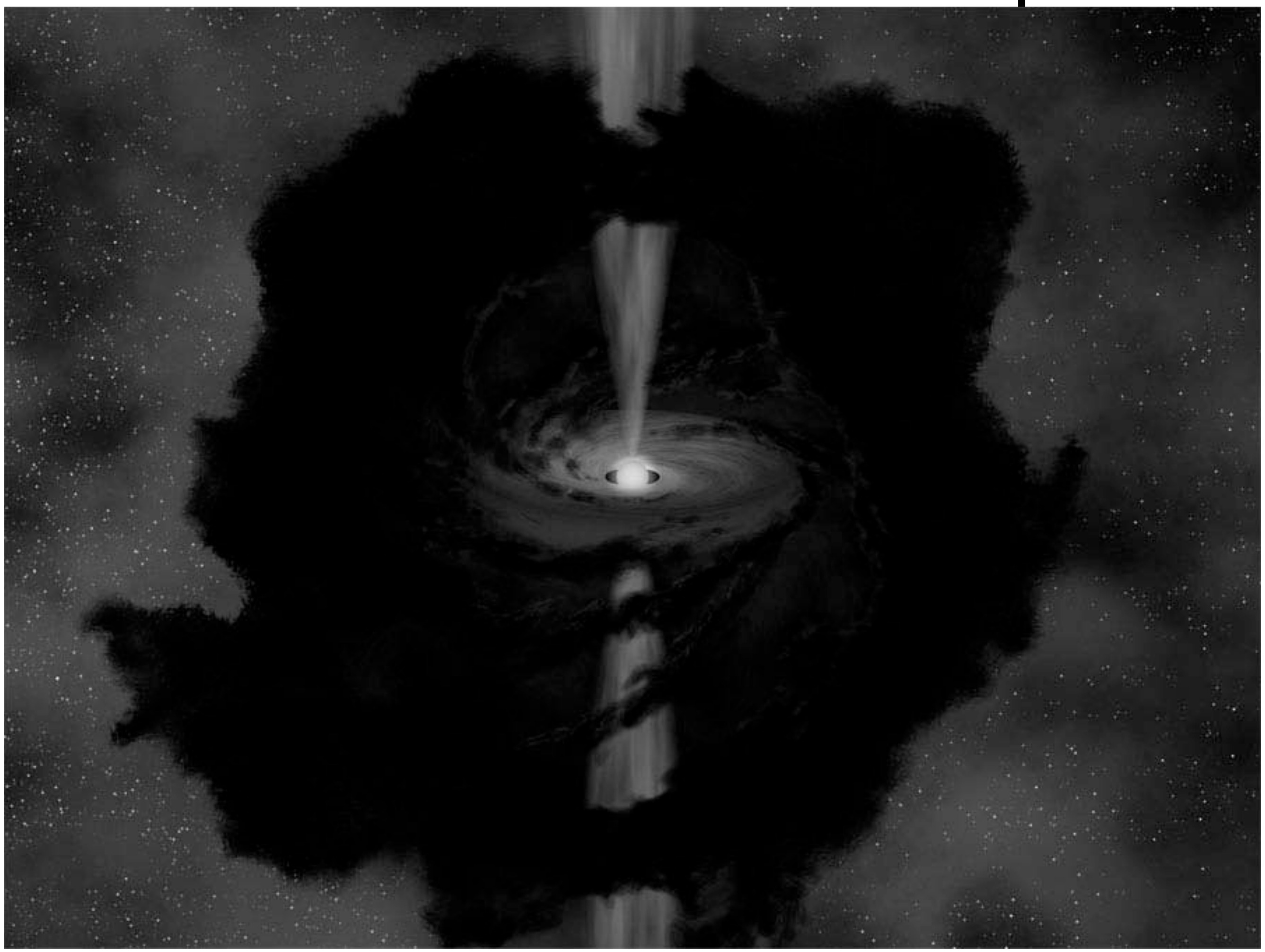

C. Lachata

Visible to Infrared Views


(2MASS)


RCW 49
234

Artist's Conception



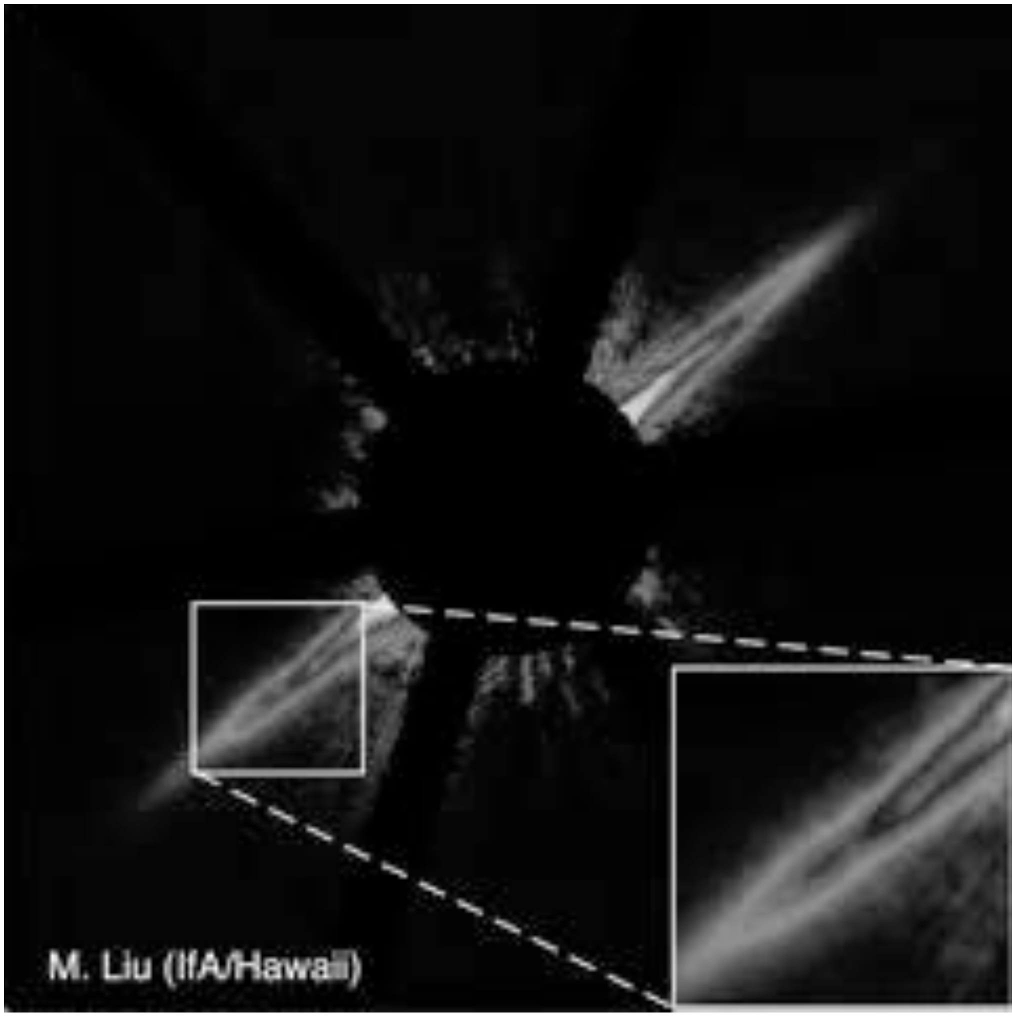
R. Hurt, SSC

Features:
Dusty envelope
Rotation
Disk
Bipolar outflow

The Protostar

- Evolution of the collapsing gas cloud
  - At first, collapsing gas stays cool
  - Dust, gas emit photons, remove energy
  - At $n \sim 10^{11}$ cm^{-3}, photons trapped
  - Gas heats up, dust destroyed, pressure rises
  - Core stops collapsing
  - The outer parts still falling in, adding mass
  - Core shrinks slowly, heats up
  - Fusion begins at $T \sim 10^7$ K
  - Protostar becomes a main-sequence star

The Disk



The Star (AU Mic) is blocked in a coronograph. Allows you to see disk. Dust in disk is heated by star and emits in infrared.

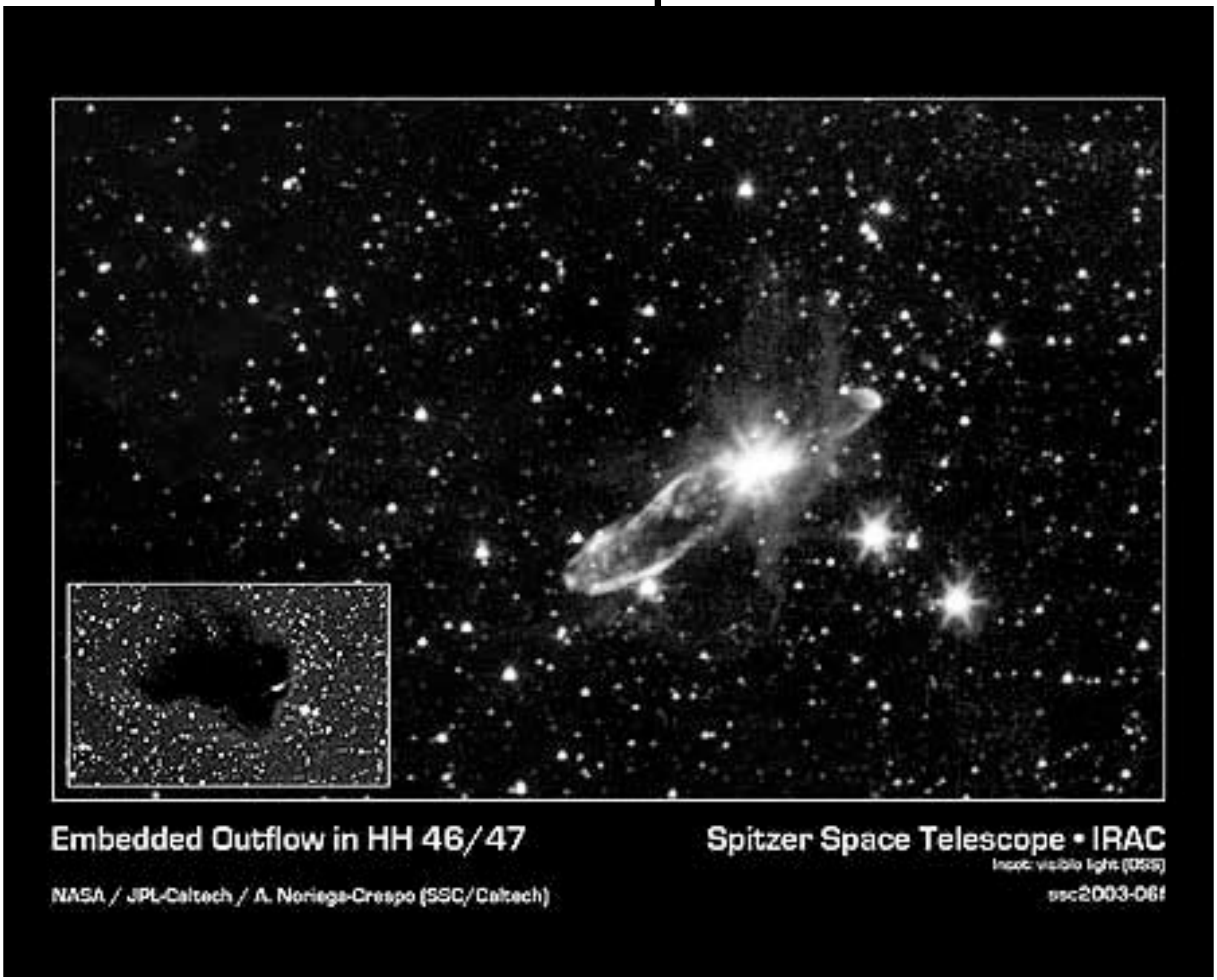
Angular Momentum

- Measure of tendency to rotate
  - J = mvr
- Angular momentum is conserved
  - J = constant
  - As gas contracts (r smaller), v increases
  - Faster rotation resists collapse
  - Gas settles into rotating disk
  - Protostar adds mass through the disk

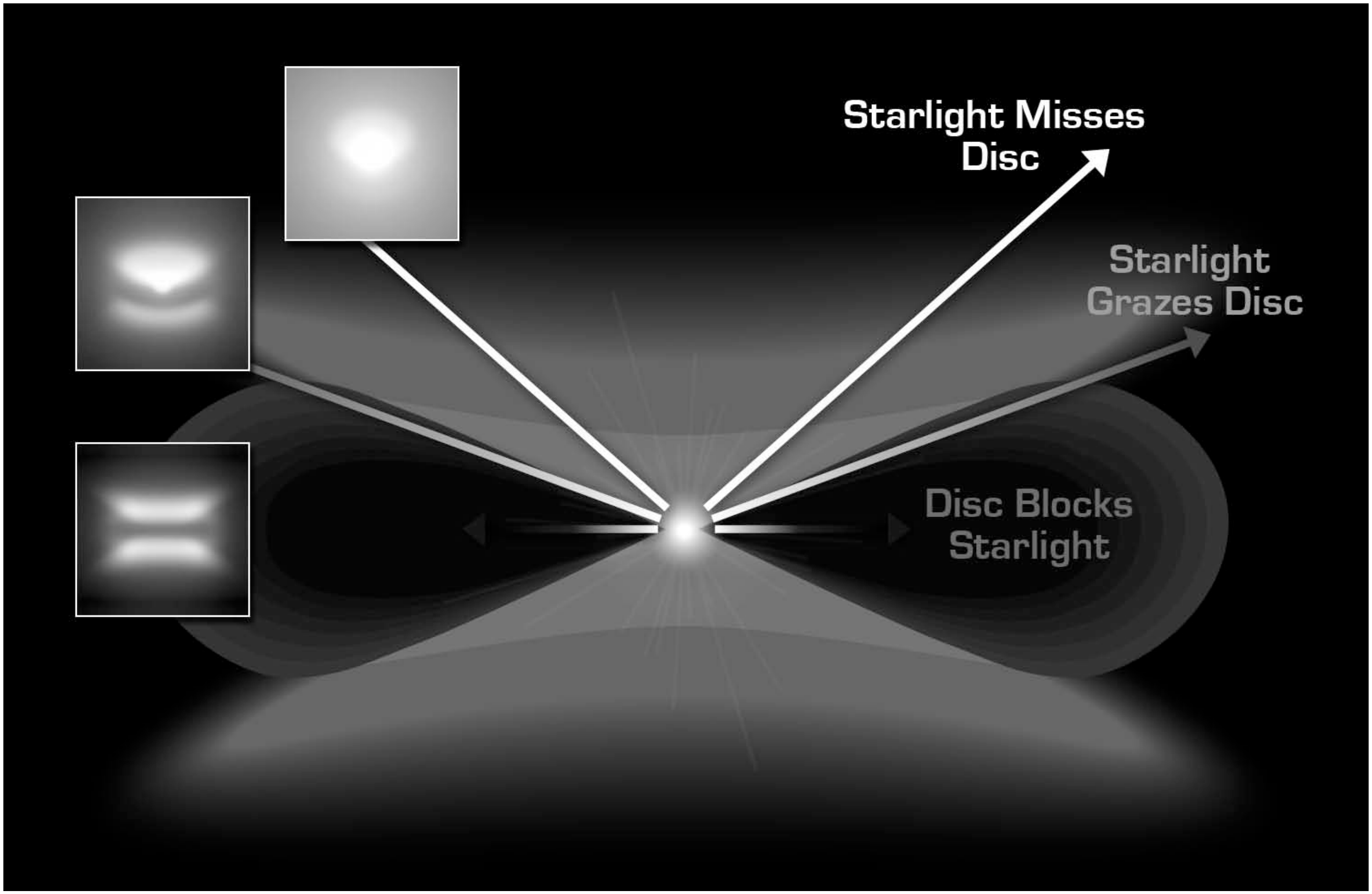
The Wind

- Accretion from disk will spin up the star
  - Star would break apart if spins too fast
- Angular momentum must be carried off
- The star-disk interaction creates a wind
- The wind carries mass to large distances
  - $J = mvr$, small amount of m at very large r
  - Allows star to avoid rotating too fast
- Wind turns into bipolar jet
  - Sweeps out cavity

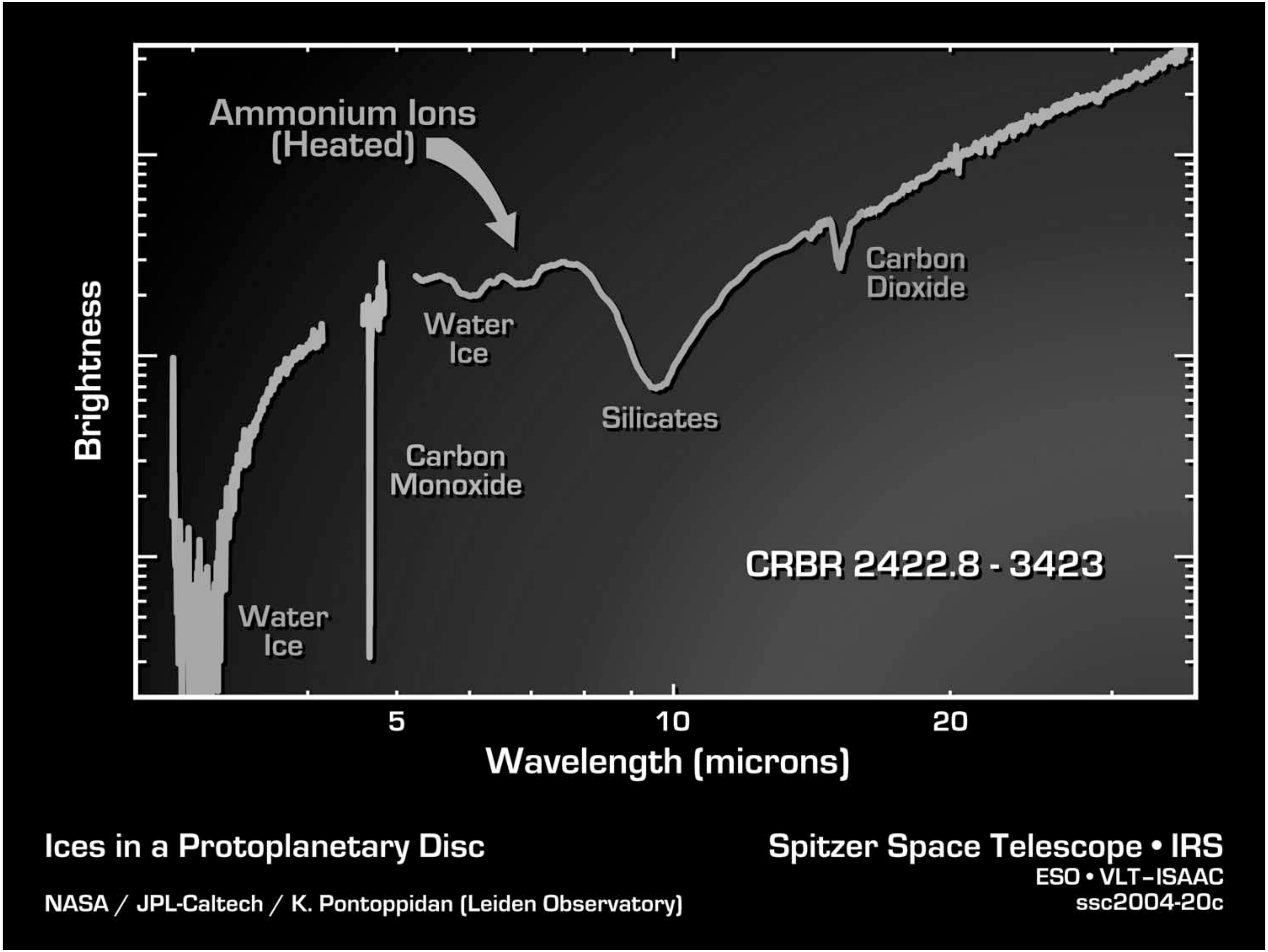
The Bipolar Jet



Studying the Disk



Robert Hurt, SSC


Ammonium Ions (Heated)
Brightness
Water Ice
Carbon Monoxide
Water Ice
Silicates
Carbon Dioxide
CRBR 2422.8 - 3423
5
10
20
Wavelength (microns)
Ices in a Protoplanetary Disc
NASA / JPL-Caltech / K. Pontoppidan (Leiden Observatory)
Spitzer Space Telescope • IRS
ESO • VLT–ISAAC
ssc2004-20c